

MicroChem 5 Results

These are what you should have seen.

	KI potassium iodide	KOH potassium hydroxide	Na ₂ C ₂ O ₄ sodium oxalate	Na ₂ SO ₄ sodium sulfate
Ca(NO₃)₂ calcium nitrate	NR	#1 - *faint white precipitate	#2 - white precipitate	NR
Cu(NO₃)₂ copper(II) nitrate	#3 - brown precipitate	#4 - blue precipitate	#5 - blue precipitate	NR
Pb(NO₃)₂ lead(II) nitrate	#6 - yellow precipitate	#7 - white precipitate	#8 - white precipitate	#9 - white precipitate
Zn(NO₃)₂ zinc nitrate	NR	#10 - milky white precipitate	#11 - *faint white precipitate	NR

NR - no reaction

*Note: This reaction is very slow, so it is possible you did not see it change in the time given.

You are to write **balanced** equations for each reaction that resulted in a precipitate (solid).

Compounds with Na⁺, K⁺, and NO₃⁻ are soluble, so they will **NOT** result in a solid precipitate.

This fact lets you determine from the results which product is the solid you saw or should have seen.

The reactions are written out below, but you need to balance them and make sure you understand how to write them when all you originally had was the reactants. The second handout sent gives you the answers & details.

