Answers to Naming Compounds Worksheet II

This is here to help you understand the naming.

Formula:	***Type of Compound:***	**Name:**
1. $FeSO_3$	*Transitional Polyatomic Ion*	1. iron (II) sulfite
2. $(NH_4)_3PO_4$	*Polyatomic Ion*	2. ammonium phosphate
3. $AlBr_3$	*Ionic*	3. aluminum bromide
4. $NaCO_3$	*Polyatomic Ion*	4. sodium carbonate
5. KNO_3	*Polyatomic Ion*	5. potassium nitrate
6. $Al_2(SO_4)_3$	*Polyatomic Ion*	6. aluminum sulfate
7. MnO	*Transitional Ion*	7. manganese(II) oxide
8. Cu_2O	*Transitional Ion*	8. copper (I) oxide
9. $CrCl_3$	*Transitional Ion*	9. chromium(III) chloride
10. N_2O_5	*Covalent Compound*	10. dinitrogen pentoxide
11. PCl_3	*Covalent Compound*	11. phosphorus trichloride
12. P_2O_5	*Covalent Compound*	12. diphosphorus pentoxide

Remember there are exceptions to the naming that are discussed in your handout.
The whole purpose of the Roman numerals is to indicate what charge the first element has. The Roman numerals will ALWAYS denote a positive charge and will ALWAYS be associated with the first element being listed, because in an ion, the first element is positive (cation) and the second element is negative (anion).
Elements in the 3A to the 6A columns have the ability to have different charges. When it is listed first, it is a positive ion, so it will use Roman numerals to represent its charge.